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counting the s-electrons with the p-electrons as outer electrons). Misunderstanding electron configuration for ions. Confusing the terms 'orbital' and 'energy level'. Dot-and-cross diagrams Knowing the general rule for individual atoms. Simple ionic compounds e.g. NaCl. Checking the total outer electrons after bonding – both ionic and

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the two atoms bonded together and the shared pair of electrons. This example clearly shows the refinement of the GCSE ideas required at AS level. Clearly stating that the bond is the electrostatic attraction between nuclei and a shared pair of electrons, this example scored 2. 4 GCE Chemistry 8CH0 01

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Bsix protons, seven neutrons and six electrons Csix protons, seven neutrons and five electrons Dseven protons, six neutrons and six electrons (c) (i) State what is meant by the term relative atomic mass. (2)

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Question number Answer Mark 4 The only correct answer is D (3p subshell 6, third quantum shell 18) A is incorrect because 2 is the number of electrons in a 3p orbital and the 3d electrons have been omitted from the third quantum shell B is incorrect because 2 is the number of electrons in a 3p orbital C is incorrect because the 3d electrons have been omitted from the third quantum shell

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and neutrons, surrounded by electrons, orbiting in shells 1.10 recall the relative mass and relative charge of a proton, neutron and electron 1.11 understand the terms atomic number, mass number, isotopes and relative

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