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terms of the  
complex reaction  
network notation  
where the rate  
for the  $j$ th

species is given

$$: N \quad k_{j1} y_1 \quad k_{j2}$$

$$y_2 \quad \dots \quad k_{lj} y_j$$

$$\dots \quad \frac{dy_j}{dt}$$

(1) A solution

is assumed of

the form:  $y_j$

$Y_j^0 e^{(2)}$  where

$Y_j^0$  and are

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determined.

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ARRHENIUS

EQUATION The

main assumption

behind the

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Arrhenius  
expression is  
that  $k = A e^{-E_a/RT}$  ( )? ( )?  
( ). This is an  
approximation,  
but it works  
quite well. The  
rate coefficient  
is the term that  
is a function of  
temperature but  
may also depend

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definition of the

activity

of species  $i$  is:

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fugacity at the standard state, that is, 1 atm for gases and thus. 3.  $K = \frac{f_{\text{NH}_3}}{f_{\text{N}_2}^{1/2} f_{\text{H}_2}^{3/2}}$  Use of the Lewis and Randall rule gives:  
$$K = \frac{f_{\text{NH}_3}}{f_{\text{N}_2}^{1/2} f_{\text{H}_2}^{3/2}} \left[ \frac{1 \text{ atm}}{1 \text{ atm}} \right]$$

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4.8. The rate of  
the chemical  
reaction doubles  
for an increase  
of 10 K in  
absolute  
temperature from

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298 K. Calculate

E a. Ans.

4.9. The

activation

energy for the

reaction, 2

$\text{HI}(\text{g}) \rightarrow \text{H}_2 + \text{I}$

2 (g) is 209.5 k

J mol<sup>-1</sup> at 581

K. Calculate the

fraction of

molecules of

reactants having

energy equal to

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or greater than  
activation  
energy?

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kinetics, also  
known as  
reaction  
kinetics, is the  
branch of  
physical  
chemistry that  
is concerned

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with  
understanding  
the rates of  
chemical

reactions. It is  
to be contrasted  
with  
thermodynamics,  
which deals with  
the direction in  
which a process  
occurs but in  
itself tells  
nothing about

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of how

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influence the

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to increasingly  
more difficult  
challenges in  
reactor design.

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third of the  
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reaction  
kinetics,  
setting the  
stage for the

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subsequent  
treatment of  
reactors  
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